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**Making Peace with P-chem—in Chemistry**

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$\ln K = (\Delta S/R) - (\Delta H/RT)$  The equilibrium for the dissociation of one mole of iodine vapour is 3.94 x10<sup>-4</sup> at 900K and the standard enthalpy is 154 kJmol<sup>-1</sup>. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant.  $\ln (K_2/K_1) = - \Delta H/R (1/T_2 - 1/T_1)$  , ANS = 6.7 x 10<sup>-2</sup> at 1200K.

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Physical chemistry is the study of how matter behaves on a molecular and atomic level and how chemical reactions occur. Based on their analyses, physical chemists may develop new theories, such as how complex structures are formed. Physical chemists often work closely with materials scientists to research and develop potential uses for new materials.

**Physical Chemistry—American Chemical Society**

Preparing for Your ACS Examination in Physical Chemistry: The Official Guide by American Chemical Society published by Chem Ed Exams (2009)

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Chemistry at Home The ACS Education Division is making educational resources available by topic to aid parents and teachers in student enrichment during this time of distance learning. We have bundled materials from the Reactions Video series, C&EN, our magazines within the Education Division and our portfolio of hands-on activities for students in grades K–12 and beyond.

**Chemistry Education Resources—American Chemical Society**

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**DEPARTMENT OF CHEMISTRY AND BIOCHEMISTRY**

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**Organic Chemistry Study Guide—NetSuite**

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