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 $\Delta H_{\text{vap}} = 8.31 \text{ J K mol}^{-1}$  (from IB Data booklet) Use  $PV = nRT$  (from IB Data booklet)  $V = \frac{nRT}{P} = \frac{1.50 \times 8.31 \times 298.15}{1.08 \times 1.01 \times 10^5} = 30.034 \text{ dm}^3$   $m = 34.1 \text{ dm}^3 \times 3[4]$  Award [4] for correct final answer. Award [3] (max) for  $0.0341 \text{ dm}^3$  or  $22.7 \text{ dm}^3$ . 3Award [3] (max) for  $34.4 \text{ dm}^3$ . 3Award [2] (max) for  $22.9 \text{ dm}^3$ . 3Award [2] (max) for  $0.0227 \text{ dm}^3$ .

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