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[6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are](#)

[equal. 15. A mole ratio is a ratio between the numbers of moles](#)

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[alanle274. Chemistry Chapter 11 Stoichiometry. stoichiometry. mole ratio. limiting reactant. excess reactant. the study of quantitative relationships between the amounts of.... in a balanced equation, the ratio between the number of moles.... a reactant that is totally consumed during a chemical reaction....](#)

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[Defining Stoichiometry pages 368–372 Practice Problems pages 371–372 1. Interpret the following balanced chemical equa-tions in terms of particles,](#)

[moles, and mass. Show that the law of conservation of mass is](#)

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[equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the](#)

[amounts of products formed by a chemical reaction](#)

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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